1. If an atom gains 1 electron, what is its new charge?

2. An atom that becomes charged is now called an _______.

3. The number of _______ in the nucleus determines the element.

4. The charge on an electron is _______.

5. Atoms with the same number of p+ but different number of n° are called _______.

6. _______ is the standard unit for atomic mass.

7. If an atom has 2 e- in its outer shell, what is its valence?

8. The mass of 1 proton is _______.

9. The charge on a proton is _______.

10. If the atoms are all the same in a single substance, then it is called an _______.

11. The charge of a neutron is _______.

12. The mass of a neutron is _______.

13. If an atom has 6 p+ and 9 e-, the charge on the atom is _______.

14. The nucleus of an atom has a _______ charge.

15. If an atom has 5 e- in its outer shell, then its valence is _______.

16. The atom in question 15 belongs to group _______.

17. Electrons circle the nucleus in energy levels called shells. How many electrons can each shell hold? K _____, L _____, M _____, N _____

18. If an atom has 10 p+ and 8 e-, the charge on the atom is _______.

19. The name of the element referred to in question 18 is _______.

20. Could the element in question 18 have a charge? Explain.
21. The smallest part of an element that can exist as an element, and still maintain the property of that element is called an _________.

22. The atomic mass of an atom with 3 p+ and 4 n° is ______.

23. Potassium 39 has 20 n°. This means that is also has _______ p+ and _______ e-.

24. The charge on the atom in question 23 is _________.

25. Draw the Bohr model for Carbon 14, the isotope of Carbon.

26. Draw the Bohr model for neon (Ne).

27. The type of bonding which involves electrostatic attraction of e- in the outer shell and the transfer of electrons is called an _____________ bond.

28. The type of bonding that involved the sharing of electrons is called a _____________ bond.

Use the formula below to answer the next few questions:

\[3\text{K}_2\text{SO}_4\]

29. There are ______ total molecules present.

30. There are ______ atoms of potassium present.

31. There are ______ atoms of sulfur present.

32. There are ______ atoms of oxygen present.

33. The mass of an e- is ________________.